

# IRON Test Paper

## for the rapid determination of ferrous and ferric iron

### Color reaction:

The presence of iron is indicated by a reddish-brown spot on yellowish-white background.

### Method of application:

Apply a drop of the weakly mineral acid test solution (pH 1–2) to the test paper. The presence of Fe ions is indicated by the appearance of a reddish-brown spot. Very high concentrations of iron or strongly acid solutions result in a bluish discoloration in the center of the reaction spot. These discolorations turn reddish-brown upon treatment with 10 % sodium hydroxide solution.

The yellow border around the reaction spot has no significance in the determination of iron.

Limit of sensitivity: 10 mg/L Iron ( $\text{Fe}^{2+/3+}$ )

### Interferences:

Vanadium interferes but only if it is present as cation. Vanadates do not interfere. Vanadium cations result in a bluish spot which resists treatment with caustic alkali solution. In the presence of Vanadium and Iron, mixed bluish-red colors appear. This interference cannot be eliminated.

Manganese results in the formation of a blue ring outside the specific reaction spot, which disappears upon treatment with 10 % sodium hydroxide solution.

Molybdate, in larger amounts, results in a yellow ring outside the specific reactions spot for iron, which resists treatment with sodium hydroxide solution.

Titanium compounds result in a yellow spot, which disappears upon treatment with sodium hydroxide solution.

### Notes:

For the selective determination of Fe(II) our Dipyridyl paper is recommended. The IRON-Test Paper also reacts with metallic iron when moistened slightly with distilled water and applied with pressure for about one minute to a metallic surface.

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